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Study Guide. STUDY.  
PLAY. Anion. An ion  
with a negative charge.  
The charge for an  
Anion is written as an  
number followed by a  
minus sign. Cation. An  
ion with a positive  
charge. The charge for  
a cation is written as a  
number followed by a  
plus sign. Chemical  
formula.

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Chapter 7 Chemistry Study Guide. STUDY. PLAY. Explain why most carbon compounds are classified as organic compounds. Carbon compounds controls what happens in the cell of living organism and play key roles in the chemistry of living organisms. Why is lead no longer used in paints for plumbing pipes.

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Good Chapter 7 Study  
Guide: Chemistry. (7.1)

Ions Key Term. To find  
the number of valence  
electrons in an atom of  
a representative

element, simply look at  
its group number. (7.1)

Ions Key Term cont. A  
positively charged ion,  
or cation, is produced  
when an atom loses  
one or more valence  
electrons. (7.1) Ions

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Key Term cont...

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Chapter 7 Chemistry  
Study Guide. STUDY.  
PLAY. valence

electrons. electrons in  
the highest occupied  
energy level of an  
element's atoms.

Group Number, 1A= 1,  
2A=2, 3A=3, 4A=4,  
5A=5, 6A=6, 7A=7,  
8A=8. How can you  
determine the number



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of valence electrons an atom has from its position on the periodic table? Name by that

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Chapter 7 Energy and  
Chemical Reactions -  
Mark Bishop 84 Study  
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Introduction to  
Chemistry Section  
Goals and  
Introductions Section  
7.1 Energy Goals To

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introduce the terms energy, kinetic energy, and potential energy. To introduce the Law of Conservation of Energy.

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### CHAPTER 7 REVIEW

Chemical Formulas and  
Chemical Compounds

### SECTION 1 SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.

c In a Stock system  
name such as iron(III)  
sulfate, the Roman  
numeral tells us (a)  
how many atoms of Fe  
are in one formula unit.

(b) how many sulfate

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ions can be attached to the iron atom. (c) the charge on each Fe ion.

## **7 Chemical Formulas and Chemical Compounds**

Chemistry Chapter 7 Study Guide. the force that holds two atoms together; may form by the attraction of a positive ion for a negative ion or by sharing electrons. the electrostatic force that holds oppositely

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charged particles  
together in an ionic  
compound.

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Chapter 7 Energy and  
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Review Skills 7.1

Energy Kinetic Energy

Potential Energy Units

of Energy Kinetic

Energy and Heat

Radiant Energy 7.2

Chemical Changes and

Energy 7.3 Ozone:

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Pollutant and Protector  
Removal of UV

Radiation by Oxygen  
and Ozone Molecules

The Natural

Destruction of Ozone

7.4

Chlorofluorocarbons: A  
Chemical Success

Story Gone Wrong

CFCs and the Ozone  
Layer.

## **Chapter 7 Energy and Chemical Reactions**

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Ch. 7-8 Chapter 7 1. Why is the quantum-mechanical model of the atom important for understanding chemistry? 2. What is light and how fast does it travel in a vacuum?

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4 Study Guide for An Introduction to Chemistry Memorize the following. (Be sure to check with your instructor to determine how much you are expected to know of the following.) SI base units (Table 1.1) The table below contains the four base units that you should know now.

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