

Chapter 7 Chemistry

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Chapter 7 Chemistry

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7.1 Introduction: Recall from Chapter 1 that solutions are defined as homogeneous mixtures that are mixed so thoroughly that neither component can be observed independently of the other.

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Solutions are all around us. Air, for example, is a solution. If you live near a lake, a river, or an ocean, that body of water is not pure H₂O but most probably a solution.

CH104: Chapter 7 - Solutions - Chemistry

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For example, the density of water at its critical point ($T = 374^{\circ}\text{C}$, $P = 217.7 \text{ atm}$) is 0.32 g/mL , about one-third that of liquid water at room temperature but much greater than that of water vapor under most conditions. The transition between a liquid/gas mixture and a supercritical phase is demonstrated for a sample of chlorine in Figure 7.6.1 ...

Chapter 7.6: Critical Temperature and Pressure - Chemistry ...

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